
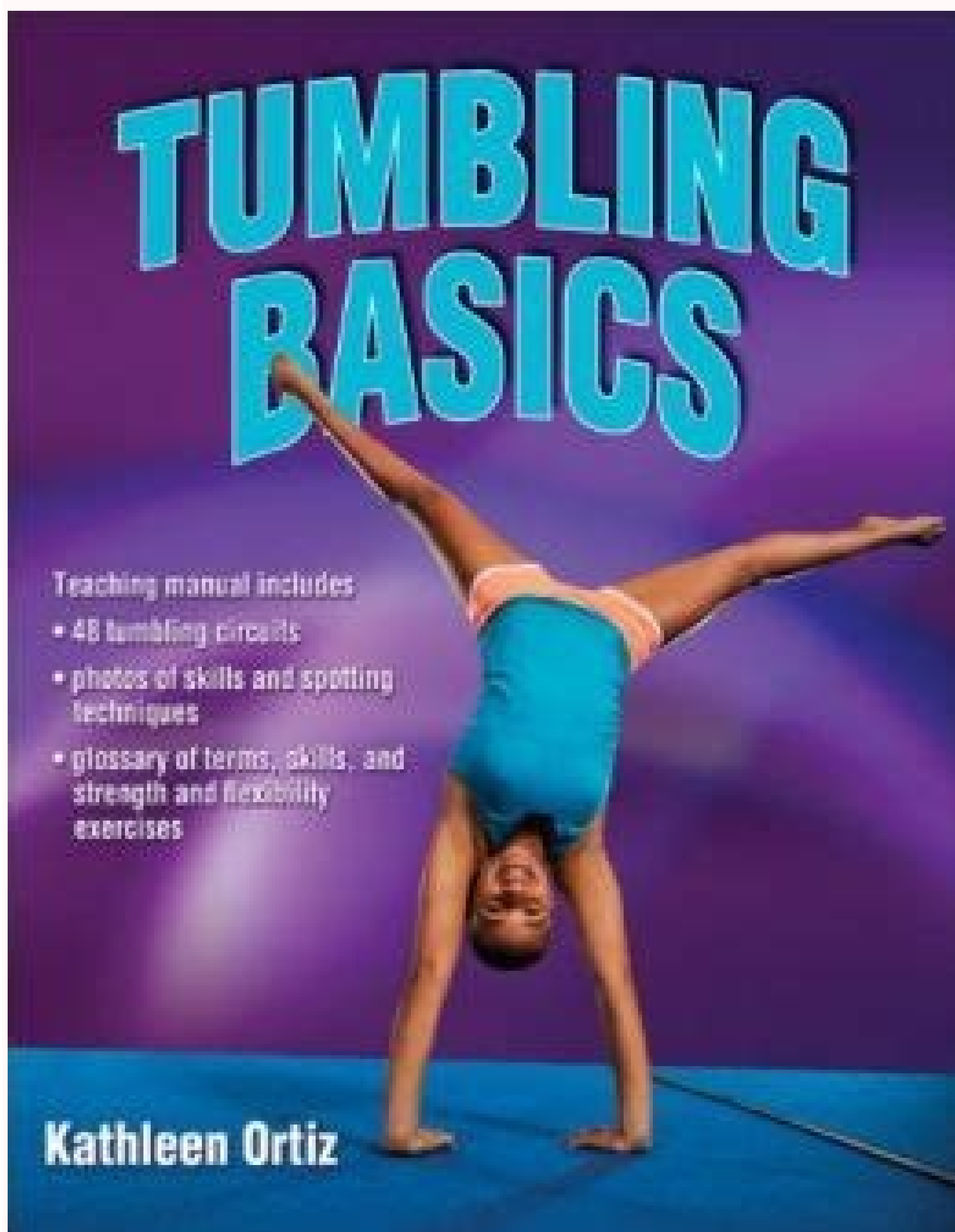


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# 5 Atoms and Molecules

## No Prerequisites

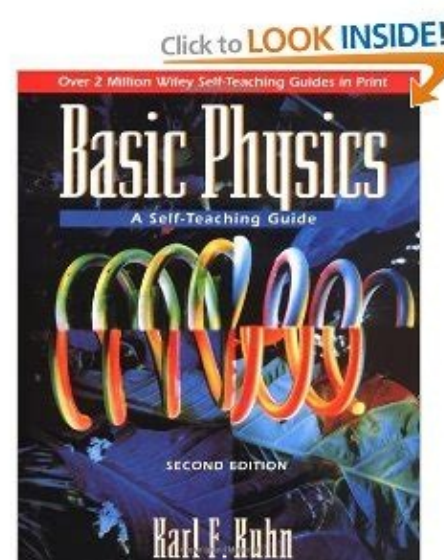
Ancient humans must have looked at the countless substances in their everyday experience and wondered if those things were perhaps made up of a smaller number of more fundamental, elementary substances. Today we know that this is indeed true—the entire physical world is made up of a relatively few substances, and these exist as atoms and molecules. The basic nature of atoms and molecules affects almost every aspect of physics. In this chapter we take a trip to the unimaginably small world of atoms.

## OBJECTIVES

After completing this chapter, you will be able to

- differentiate between elements and compounds;
- specify the total number of elements now known;
- differentiate between atoms and molecules;
- relate atoms and molecules to elements and compounds;
- give an example of the application of the Law of Definite Proportions;
- identify the relative number and type of atoms in a compound, given its formula and the periodic table;
- label the locations of the atomic particles on a drawing of the Bohr model of the atom;
- identify the component of an atom that determines its chemical properties;
- define atomic mass and atomic number;
- interpret portions of the periodic table;
- specify, using an example, the importance of the number of electrons in the outer orbit of an atom;

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